

Conformational equilibria involving 2-amino-1,3-dioxans: steric control of the anomeric effect

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Dedicated to Ossie Tee, with respect and affection

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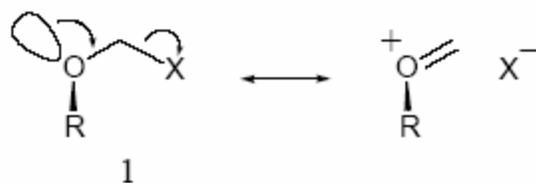
Abstract

Although *N,N*,1-trimethylcyclohexamine exists in two equilibrating conformations, with the NMe₂ group axial and equatorial in comparable amounts, in *N,N*,2-trimethyl-2-amino-1,3-dioxane the conformation with the amine equatorial could not be detected; evidence for an anomeric effect for the NMe₂ group of >3 kJ mol⁻¹. When the NMe₂ group is axial in the dioxane, there can be no *exo*-anomeric effect (this would mean placing a methyl group over the ring), but the *endo*-anomeric effect is unaffected. Calculations show that when the NMe₂ group is equatorial, the rotamer lowest in energy is one in which an *exo*-anomeric effect is possible, despite the steric demands involved. Thus the *endo*-anomeric effect in *N,N*,2-trimethyl-2-amino-1,3-dioxane is greater than the *exo*-anomeric effect, even though amine nitrogen is generally a better electron-pair donor than oxygen.

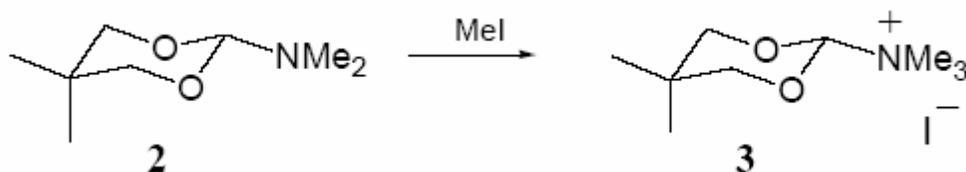
Keywords: Anomeric effect, steric control, 2-amino-1,3-dioxans, conformational equilibria

Introduction

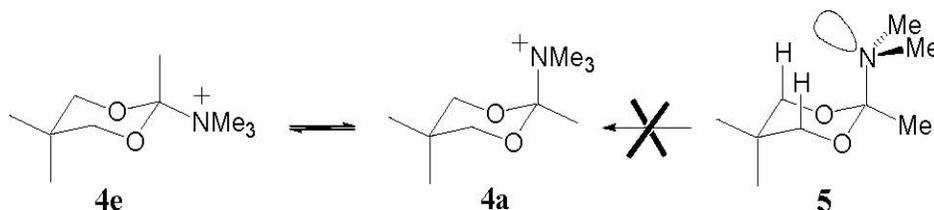
The generalized anomeric effect¹ refers to the predilection for *gauche* conformations of systems such as **1**, with two heteroatoms on the same carbon. Any conformational effect is the result of a combination of steric, electronic and electrostatic influences, but there is substantial theoretical support² for charge demoralization derived from hyperconjugation (arrows in **1**) as the primary cause of the anomeric effect. The effect is stronger for more electronegative groups X, and for better electron-pair donors (which may be lone pairs, as in **1**, or even σ -bonding electrons¹), as the n_O- σ^* C-X interaction (arrows in **1**) becomes more important. We report results with an O—C—N system, which shows an unexpectedly large anomeric effect.



As part of our investigation of the putative reverse anomeric effect^{3,4} we were interested in systems **3** and **4**. **3** is readily made by alkylation of amide acetal **2** with methyl iodide at room temperature.⁵



The equatorial protons on C4 and C6 of **2** resonate downfield of their geminal proton partners (as is normal for THPs and cyclohexanes), but in the quaternary ammonium compound **3** the order is reversed: the axial protons H(4) and H(6) appearing downfield of their equatorial counterparts. The same effect [$\delta_{\text{H}(4_{\text{ax}})} > \delta_{\text{H}(4_{\text{eq}})}$] has been noted before in 2-alkoxy-1,3-dioxanes, where it was ascribed to the axial preference of the alkoxy group,^{6,7} but nOe experiments confirmed that in **3** the $^+\text{NMe}_3$ group is indeed equatorial. Intending to characterize the conformation with the $^+\text{NMe}_3$ group axial we set out to make compound **4**, with a sterically balancing methyl group on C2.



Amide acetal **5** was readily made by reacting *N,N*-dimethylacetamide dimethylacetal with the appropriate diol, but proved impossible to methylate using methyl iodide, or even trimethyloxonium tetrafluoroborate. nOe experiments revealed that **5** actually exists in the conformation with the NMe_2 group axial. (The ^1H NMR shifts of the axial and equatorial protons in **5** are 3.82 and 3.10 respectively, confirming the preferred conformation.) We presume that **5** is difficult to methylate because the axial NMe_2 group has its lone pair situated directly over the dioxane ring, thus sterically inhibiting its nucleophilicity. Allinger and Graham have shown that the dimethylamino group in a cyclohexane ring is alkylated 50 times more slowly when axial than when equatorial.⁸ In a dioxane the transition state would be yet more crowded, because C—O bonds are shorter than C—C.

Because amine nitrogen is less electronegative than oxygen, groups such as NH_2 show the

usual steric preference for an equatorial configuration and only a small anomeric effect. A-values for the NH₂, NHMe and NMe₂ groups are 1.23, 1.29 and 1.53, respectively,⁹ all smaller than the value of 1.74 for CH₃, so that **5** is expected to prefer the conformation with the dimethylamino-group axial by a small margin. The fact that no trace could be detected of the conformer of amide acetal **5** with the dimethylamino-group equatorial was thus unexpected, and provided an opportunity to estimate the size of the anomeric effect associated with the dimethylamino group.

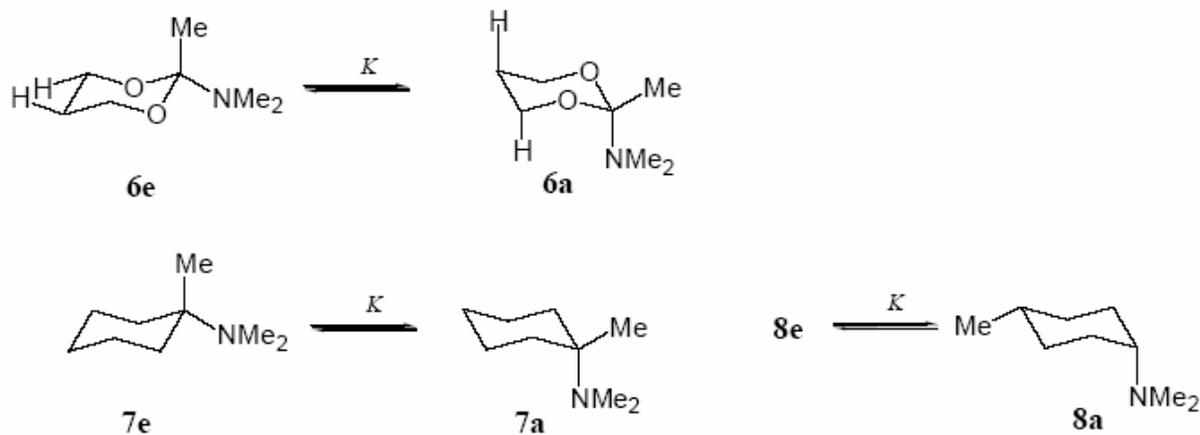
Care must be taken when using A-values alone when trying to distinguish between steric and electronic effects. For example, 2-methylaminotetrahydropyran (see discussion, below) shows a stronger preference than methylaminocyclohexane for the equatorial conformation, leading Booth and Khedhair¹⁰ to conclude that the NHMe group has a small reverse anomeric effect. However, this conclusion takes no account of the stronger 1,3-diaxial interactions present in the dioxane ring because of the shorter C–O bonds. (The A-value for the methyl group at room temperature is 1.74 kcal mol⁻¹, but Eliel and Knoeber determined ΔG° for the ring inversion of 2-methyl-1,3-dioxane as >3.5

kcal mol⁻¹.¹¹ Making no allowance for the stronger interactions in the dioxane, these data would suggest a strong reverse anomeric effect for the methyl group!)

The method of Franck allows predictions of the effective sizes of substituents on saturated heterocyclic rings by comparing values of ΔG° for the ring-inversion of substituted systems with A-values for the same equilibria in similarly substituted cyclohexanes.¹² The A-value of the methyl group, for example, is compared with the standard conformational free energy change for the ring inversion of the methyl-substituted dioxane ring, and the same proportionality factor applied to A-values of other groups. Franck's original paper compared the sizes of 'inert' groups like methyl, in cyclohexyl and tetrahydropyranyl systems, but the method is simply extended to 1,3-dioxane and other saturated heterocyclic systems.¹³ With steric corrections taken into account, Franck concluded that NMe₂ and NHMe groups actually show a slight anomeric effect.¹²

Results and Discussion

We studied amide acetal **6** to provide some insight into the magnitude of the anomeric effect of the dimethylamino group, by comparing ΔG° for ring inversion with that of its cyclohexyl analogue **7**. As in our original system **5** the geminal methyl group acts as a steric counterbalance to the NMe₂ group. In the analysis, we assume, with Franck, that both groups experience the same degree of steric compression on going from the cyclohexane to the dioxane ring. If there is no anomeric effect in dioxane **6** the equilibrium constants for the ring inversion of **6** and **7** (Scheme 1) should be the same.



Scheme 1

We can compare our data for the ring inversion of amine **7** with the results of Booth and Jozefowicz, who measured the equilibrium constant for the related amine **8** in order to determine the *A*-value for the dimethylamino group.¹⁴ In this system the 4-methyl group was used as a counterbalance, so the value for $-\Delta G^\circ$ was subtracted from the *A*-value for the methyl group (taken as $1.7 \text{ kcal mol}^{-1} = 7.11 \text{ kJ mol}^{-1}$) to give *A*-values for the dimethylamino group of 6.40 kJ mol^{-1} ($1.53 \text{ kcal mol}^{-1}$) in $\text{CFCl}_3:\text{CDCl}_3$ 9:1 and 5.48 kJ mol^{-1} ($1.31 \text{ kcal mol}^{-1}$) in toluene.¹⁴ In our system **7** the equilibrium constant for the ring inversion is closely similar (Table 1), and we conclude that, at least in this geminally disubstituted compound, there is little or no interaction between the two groups.

Table 1 Comparison of equilibrium constants for ring inversion of *N,N*,1- and *N,N*,4-trimethylcyclohexylamines **7** and **8**

| Compound | Solvent | Temperature / K | <i>K</i> | $-\Delta G^\circ /$ kJ mol ⁻¹ |
|-------------|--|-----------------|-----------------|---|
| 8 14 | $\text{CFCl}_3:\text{CDCl}_3$ 9:1 | 183 | 1.61 | 0.71 |
| 8 14 | toluene- <i>d</i> ₈ | 183 | 3.0 | 1.63 |
| 7 | $\text{CFCl}_3:\text{CDCl}_3$ 9:1 | 180 | 2.81 ± 0.07 | 1.55 ± 0.03 |
| 7 | $\text{CFCl}_3:\text{CD}_2\text{Cl}_2$ 9:1 | 185 | 2.82 ± 0.04 | 1.59 ± 0.03 |

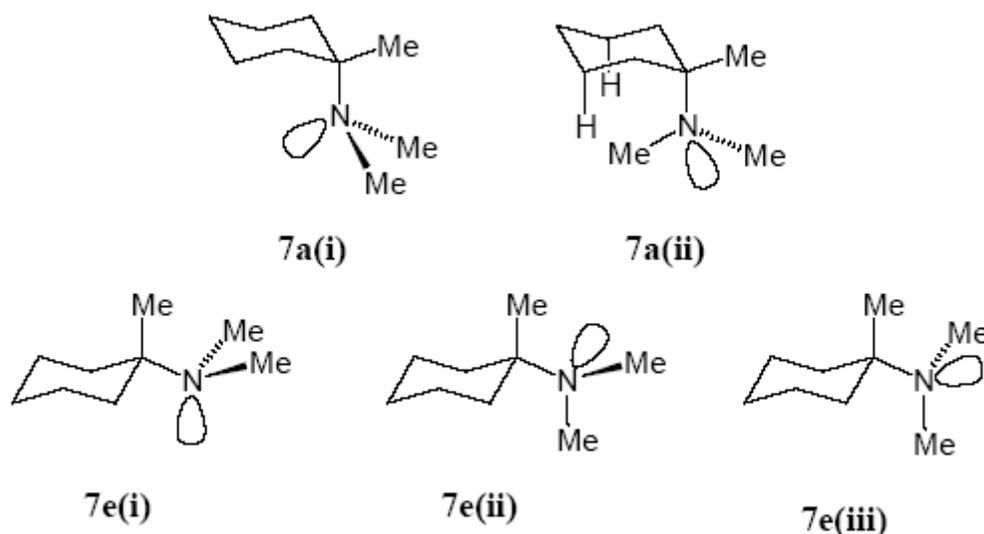
Table 2 Equilibrium constants and standard free energies for the ring inversion of **7** at different temperatures

| Temperature / K | <i>K</i> | $-\Delta G^\circ$ |
|-----------------|-----------------|-------------------|
| 190 | 2.90 ± 0.11 | 1.68 ± 0.06 |
| 185 | 2.82 ± 0.04 | 1.59 ± 0.03 |
| 180 | 2.81 ± 0.07 | 1.55 ± 0.03 |
| 175 | 2.83 ± 0.04 | 1.51 ± 0.02 |
| 170 | 3.00 ± 0.08 | 1.55 ± 0.04 |

A plot of $\ln K$ against $1/T$ gave the thermodynamic parameters as:

$$\Delta H^\circ = -0.41 \pm 0.48 \text{ kJ mol}^{-1}, \Delta S^\circ = 6.5 \pm 3 \text{ J K}^{-1} \text{ mol}^{-1}$$

We note that for the ring inversion of amine **7**, ΔH° is zero (within experimental error, Table 2) and ΔS° is positive, i.e. ΔG° is predicted to vary with temperature. This positive entropy difference for the ring inversion is unexpected. When the dimethylamino group is axial, the conformation is restricted to the single rotamer, **7a(i)**, with the lone pair over the ring – rotamers with methyl groups over the ring are much higher in energy (Scheme 2). When the group is equatorial rotamer **7e(i)** should be most favorable, though rotamers **7e(ii)** and **(iii)** should not be inaccessible in energy.



Scheme 2

Thus the simple prediction is that the conformation with the dimethylamino group *equatorial* should have the greater entropy, the opposite to what is observed. Höfner *et al.* also observed a small positive entropy change ($\Delta S^\circ = 1.8 \text{ kJ mol}^{-1}$) when the methoxy group in

methoxycyclohexane goes from equatorial to axial.¹⁵ They point out that many other contributions to the entropy change that must be considered besides rotational factors, and try and quantify some of these. Praly and Lemieux, who analysed solvent contributions to entropy changes, suggest stronger solvation of the equatorial conformer of 2-methoxytetrahydropyran to explain the positive $\Delta S^\circ_{\text{eq}\rightarrow\text{ax}}$.¹⁶

Our objective was to compare the free energy changes for the ring inversions of amine **7** and amide acetal **6**, using low temperature ¹H NMR. However, on cooling a sample of amide acetal **6** to 140 K (in 70:30 CBr₂F₂:CD₂Cl₂), no trace was observed of the conformer with the NMe₂ group equatorial. This could mean either the temperature was not low enough for the two conformers to be in slow exchange, or there was only the one conformer present. Since there was no significant change in appearance of the lowest temperature spectra, although dioxanes are usually in slow exchange well above 150 K, we conclude that the amide acetal exists effectively as a single conformer. nOe experiments tell us that the conformation present is **6a**.

The coupling constants in the ¹H NMR spectra of **6** (Table 3) are consistent with this conclusion. Were the amide acetal in rapid equilibrium with its conformer, the coupling constant between the protons on C(4) and C(5) would represent a weighted average of the coupling constants in the two conformations (see Scheme 1). The *trans*-diaxial coupling constant between 4H_{ax} and 5H_{ax} is so large (12.7 Hz, Table 3) that **6e** cannot be present in significant amounts.¹⁷

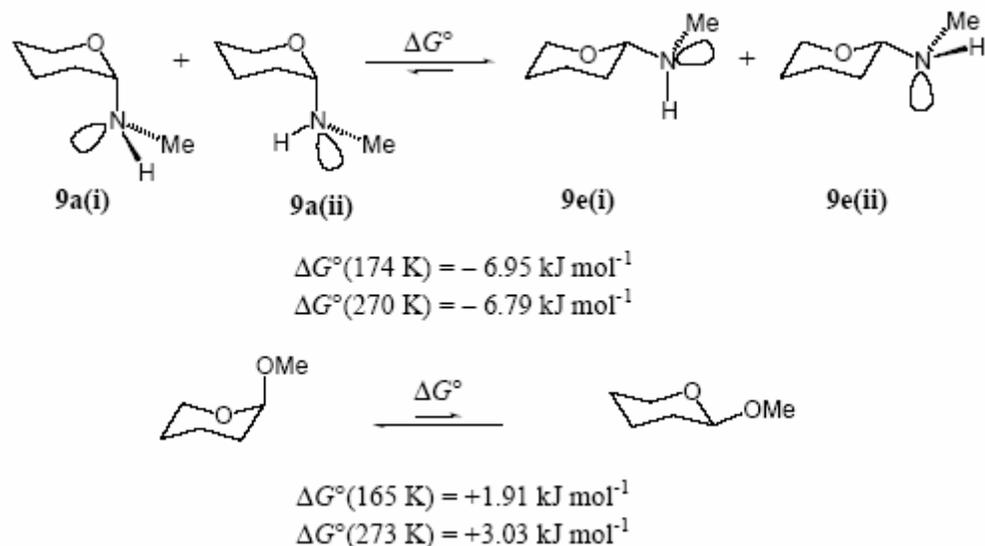
Table 3 The ¹H-¹H coupling constants in amide acetal **6**

| Coupling | Coupling constant / Hz |
|-------------------------------------|------------------------|
| ² J _{H4ax-H4eq} | -11.4 |
| ³ J _{H4ax-H5ax} | 12.7 |
| ³ J _{H4ax-H5eq} | 2.8 |
| ³ J _{H4eq-H5ax} | 5.2 |
| ³ J _{H4eq-H5eq} | 1.4 |
| ² J _{H5ax-H5eq} | -13.2 |

If the steric compression experienced by the methyl and dimethylamino groups increases by the same factor on going from the cyclohexane to the dioxane (as assumed in Franck's method), then the equilibrium constants for the ring inversions of the two compounds should be the same. At 185 K, the NMe₂ axial conformation for the cyclohexane compound is more favourable by $\Delta G^\circ = 1.59 \text{ kJ mol}^{-1}$, whereas for the dioxane analogue (assuming, conservatively, a maximum ratio, axial:equatorial 95:5 for NMe₂), the NMe₂ axial conformation **6a** is favoured by $\Delta G^\circ \geq 4.5 \text{ kJ mol}^{-1}$. Hence, the conformation of dioxane **6** with the axial NMe₂ group is at least 3 kJ mol⁻¹ more stable at 185 K than would be expected from steric factors alone.

This result can be used to shed some light on the *exo/endo* anomeric effects in the 2-aminotetrahydropyrans studied extensively by Booth *et al.*^{9,18,19} Because nitrogen is less electronegative than oxygen it is a better donor, and $\sigma^* \text{C}-\text{N}$ a weaker acceptor of electrons

than σ^*C-O . So 2-methylaminotetrahydropyran should exhibit a stronger *exo*-anomeric effect ($n_N-\sigma^*C-O$) than *endo*-anomeric effect ($n_O-\sigma^*C-O$). Thus, as shown by Booth *et al.*, in contrast to 2-methoxytetrahydropyran, 2-methylaminotetrahydropyran actually prefers the equatorial conformation (Scheme 3).¹⁹



Scheme 3

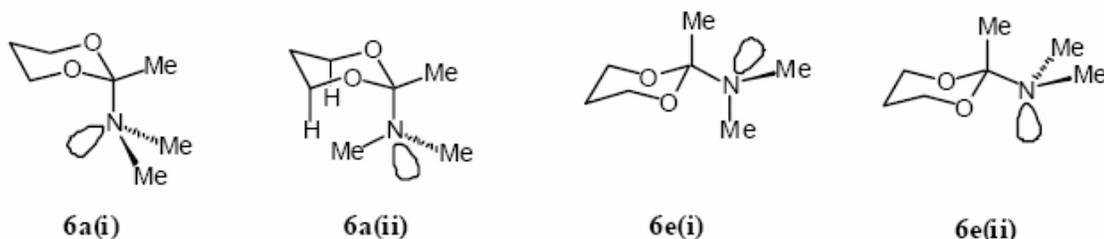
By studying homo- and hetero-nuclear coupling constants in **9e** and **9a**, these authors were able to confirm the preference for rotamers exhibiting an *exo*-anomeric effect (Scheme 3). With the NHMe group equatorial rotamer **9e(i)**, stabilised by an *exo*-anomeric effect, is preferred. With NHMe axial the preference is for rotamer **9a(ii)**: for the same reason, even though this means that the hydrogen atom must lie above the ring.

These results, together with the finding that 2-dimethylaminotetrahydropyran also prefers the conformation with the NMe₂ group equatorial,²⁰ prompted Eliel *et al.* to conclude that “in [2-dimethylamino- and 2-methylamino-tetrahydropyrans], the equatorial isomer is actually favoured since the *exo*-anomeric effect now outweighs the *endo*-anomeric effect”.²¹ Our results show that in *N,N*,2-trimethyl-2-amino-1,3-dioxane the dimethylamino group prefers the axial conformation: even when size factors are taken into account, the *exo*-anomeric effect does not outweigh the *endo*-anomeric effect.

In contrast to the *N*-methylaminotetrahydropyrans, no *exo*-anomeric effect is expected when the dimethylamino group is axial in the 2-position on a dioxane ring, because in order for the nitrogen lone pair to be *anti* to a C–O bond, a methyl group would have to lie over the ring (**6a(ii)**, Scheme 4). This is possible for a hydrogen (*cf.* **9a(i)**), but the steric demands of a methyl group are prohibitive.

The question then arises, is an *exo*-anomeric effect still possible when the NMe₂ group is equatorial? For an *exo*-anomeric effect to be possible (rotamer **6e(i)**), one of the methyl groups

on the nitrogen must align itself in an axial-like position (Scheme 4). Though sterically less favourable than rotamer **6e(ii)**, this should be possible, since the eclipsing 1,3-interactions are with lone pairs on oxygen. For guidance we have performed *ab initio* calculations on **6** and **7**.



Scheme 4

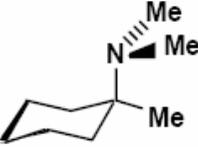
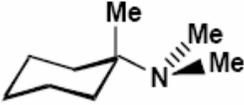
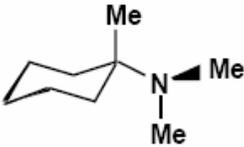
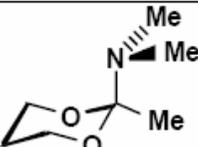
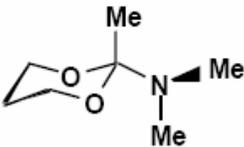
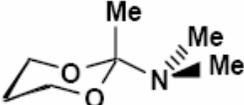
Calculations

Calculations on the energies of different conformers of **6** and **7** proved illuminating. Results using MM2 and a 6-31G* basis set are compared in Table 4. Both techniques predict the order that is observed experimentally, with the conformation with the NMe₂ group axial preferred in both cyclohexane and dioxane systems. For the dioxane, conformation **6a(i)**, with the NMe₂ group axial, is substantially lower in energy than the equatorial pair **6e(i)** and **(ii)**. However conformation **6e(i)** is lower in energy than **6e(ii)**.

This is contrary to the order observed in the cyclohexane and suggests that when the NMe₂ group is equatorial, an *exo*-anomeric effect is possible.

If a strong *exo*-anomeric effect is in fact possible when the dimethylamino group is equatorial, we might have expected to see at least some evidence of this conformation. Perhaps another reason why this is not observed is that in the dioxane, it is possible for *both* ring oxygen atoms to interact with the σ^* C–N orbital in the conformation with the NMe₂ group is axial, while the single lone pair on the nitrogen can interact with only one σ^* C–O orbital when equatorial.

Table 4 Results of theoretical calculations on conformers of **6** and **7**

| Conformation | Number | Relative Energy / kJ mol ⁻¹ | |
|---|--------|--|--------|
| | | MM2 | 6-31G* |
|  | 7a(i) | 0 | 0 |
|  | 7e(i) | 4.5 | 6.3 |
|  | 7e(ii) | 8.8 | 12.0 |
| <hr/> | | | |
|  | 6a(ii) | 0 | 0 |
|  | 6e(i) | 15.93 | 20.3 |
|  | 6e(ii) | 20.5 | 30.8 |

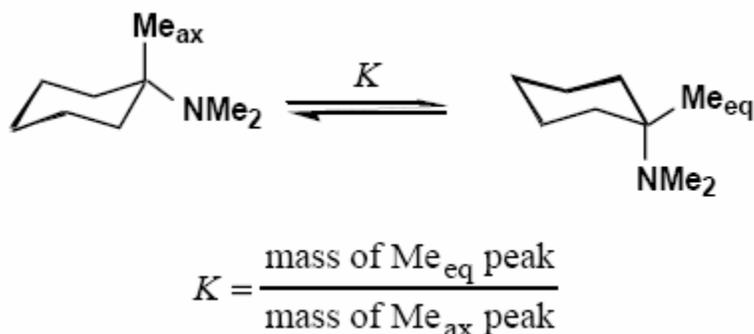
Conclusions

N,N,2-trimethyl-2-amino-1,3-dioxane **6**, like its 5,5-dimethyl derivative **5**, exists apparently exclusively in the conformation with the dimethylamino-group axial. We conclude that the *endo*-anomeric effect in this system is greater than the *exo*-anomeric effect, even though amine nitrogen is generally a better electron-pair donor than oxygen.

Experimental Section

Low-temperature ^1H NMR measurements. We determined the thermodynamic parameters for the ring inversion of amine **7** by direct measurement of the conformational equilibrium constants using low-temperature ^1H NMR spectra. At room temperature, the NMe_2 protons and the C-methyl protons appear as singlets at 2.22 ppm and 0.88 ppm respectively, and both are separated from other signals by at least 0.35 ppm. However, the peaks due to the NMe_2 protons in the two conformations are too close together (δ 2.14 ppm and 2.11 ppm) for accurate determination of their relative proportions. At 200 K, the signal due to the C-methyl group has separated into two clearly resolved peaks – the smaller at 0.86 ppm (from the conformation with the methyl group axial) and the larger at 0.76 ppm (methyl group equatorial). A series of spectra recorded at temperatures between 298 K and 190 K show progressive changes in the ^1H NMR spectra of **7**.

For the quantitative measurements, samples were made up to approximately 0.1 mol dm^{-3} in 9:1 $\text{CFCl}_3:\text{CD}_2\text{Cl}_2$. Spectra were recorded on a Bruker AM400 spectrometer in 5 mm tubes. Samples were allowed to equilibrate at 5 K intervals from 190 K to 175 K and four spectra were recorded at each temperature. The resulting spectra were plotted over the region of the methyl peaks for the two conformers and then expanded further by photocopying. The two peaks were then cut out, accurately weighed and the masses used to determine the conformational equilibrium constants using the following equation:



The four values of K for each temperature were then averaged to give the values shown in Table 2.

Experimental Section

General Procedures. All solvents were distilled before use and dried, unless otherwise indicated, by stirring over, then distilling from, a suitable drying agent: lithium aluminium hydride for tetrahydrofuran and diethyl ether; calcium hydride for dichloromethane, hexane, toluene and cyclohexane and phosphorus pentoxide for glacial acetic acid. Commercial

cyclohexanol, which contains trace amounts of cyclohexanone was purified by cooling and filtering off the pure solid cyclohexanol (freezing point 24 °C) and then distilling.

Merck Kieselgel 60 (230-400 mesh) was used for flash column chromatography. Merck Kieselgel 60F254 pre-coated plates (0.25 mm) were used for thin layer chromatography and the products were visualized with iodine vapor or potassium permanganate solution.

Melting points were measured with a Reichert hot stage microscope and are uncorrected. Infrared spectra were recorded on a Perkin-Elmer 1600 FTIR spectrophotometer. Proton and carbon NMR spectra were recorded on a Bruker AC 200, Bruker WM 250, Bruker AM 400 or a Bruker DRX 500 Fourier transform spectrometer. Mass spectra were recorded on an AEI Kratos MS30 or an AEI Kratos MS890 machine. A DS503 data system was used for high resolution analysis. Microanalyses were carried out by the staff of the University Chemical Laboratories using Carlo Erba 1106 or Perkin-Elmer 240 automatic analysers.

N,N,5,5-Tetramethyl-1,3-dioxan-2-amine (2). A solution of 2,2-dimethyl-1,3-propanediol (4 g, 39 mmol) and a crystal of *p*-toluenesulphonic acid were heated in dry toluene to 90 °C under argon, and *N,N*-dimethylformamide dimethyl acetal (5 g, 42 mmol) added drop wise. The methanol formed was distilled out as the azeotrope with toluene, the solution cooled, washed with sodium bicarbonate solution (25 mL) and brine (25 mL), dried (MgSO₄) and evaporated under reduced pressure. The residue was distilled to give the amide acetal (3.27 g, 53%), bp 56-57 °C / 7 mmHg. (lit.²¹ 68-70 °C/7mmHg); IR: ν_{\max} (Film)/cm⁻¹ 2953 and 2870 (C–H), 1111 and 1097 (C–O); ¹H NMR (200 MHz; CDCl₃) δ 4.71 (1 H, s, H₂), 3.62 (1 H, d, *J* 10, H_{4eq} and H_{6eq}), 3.47 (1 H, d, *J* 10, H_{4ax} and H_{6ax}), 2.40 (6 H, s, -N(CH₃)₂), 1.17 (3 H, s, C5-Me_{ax}) and 0.73 (3 H, s, C5-Me_{eq}); ¹³C NMR (CDCl₃) δ 108.8 (C2), 76.2 (C4 and C6), 37.7 (-N(CH₃)₂), 29.7 (C5), 22.8 (-CH₃)_{ax}) and 21.4 (-CH₃)_{eq}) (Anal. C₈H₁₇NO₂ Found: C, 60.15; H, 10.70; N, 8.75. Requires C, 60.35; H, 10.75; N, 8.80%).

N,N,N,5,5-Pentamethyl-1,3-dioxan-2-aminium iodide (3). Methyl iodide (0.68 g, 4.3 mmol) was added to a stirred solution of *N,N,5,5*-tetramethyl-1,3-dioxan-2-amine **2** (0.75 g, 4.3 mmol) in dry ether (10 mL) and the solution stirred for 36 h under argon. The precipitate formed was filtered off and washed with ether to give the quaternary amide acetal⁵ as a white powder (0.92 g, 71%); IR: ν_{\max} (KBr)/cm⁻¹ 3014, 2968, 2864 and 2738 (C–H), 1105 and 1056 (C–O); ¹H NMR (400 MHz; CD₂Cl₂) δ 6.26 (1 H, s, H₂), 3.94 (2 H, d, *J* 11, H_{4ax} and H_{6ax}), 3.81 (2 H, d, *J* 11, H_{4eq} and H_{6eq}), 3.32, (9 H, s, -N(CH₃)₃⁺), 1.19 (3 H, s, C5-(CH₃)_{ax}) and 0.82 (3 H, s, C5-(CH₃)_{eq}); ¹³C NMR (CD₂Cl₂): δ 106.6 (C2), 77.0 (C4 and C6), 49.1 (-N(CH₃)₃⁺), 30.4 (C5), 22.8 (C5-(CH₃)_A(CH₃)_B) and 20.6 (C5-(CH₃)_A(CH₃)_B); MS: *m/z* 115 (100%, M – NMe₃); Anal. C₉H₂₀NO₂I Found C, 35.75; H, 6.70; N, 4.60; I, 41.90 requires C, 35.90; H 6.70; N 4.65; I, 42.15%).

N,N,2,5,5-Pentamethyl-1,3-dioxane-2-amine (5). A solution of 2,2-dimethyl-1,3-propanediol (1.25 g, 12 mmol) and a crystal of *p*-toluenesulphonic acid in dry toluene (30 mL) was heated to 90 °C under argon and *N,N*-dimethylacetamide dimethyl acetal (2 g, 15 mmol) added dropwise. The methanol-toluene azeotrope formed was distilled out and the solution stirred for 3 h at

90 °C. The solution was cooled, evaporated under reduced pressure and the residue distilled to give the *amide acetal* (0.79 g, 38 %) as a colourless oil, bp 44-46 °C/6 mmHg; IR: $\nu_{\max}(\text{Film})/\text{cm}^{-1}$ 2997, 2945 and 2867 (C–H), 2821 and 2783 (NMe₂ C–H), 1372 (Nme₂), 1225, 1200, 1184, 1134, 1079 and 1037 (C–O); ¹H NMR (200 MHz; CDCl₃) 3.82 (2 H, d, *J* 10, H_{4ax} and H_{6ax}), 3.10 (2 H, d, *J* 10, H_{4eq} and H_{6eq}), 2.20 (6 H, s, -N(CH₃)₂), 1.20 (3 H, s, C2-CH₃), 1.16 (3 H, s, (C5-(CH₃)_{ax}) and 0.72 (3 H, s, (C5-(CH₃)_{eq}); ¹³C NMR (CDCl₃) 105.9 (C2), 68.9 (C4 and C6), 37.5 (-N(CH₃)₂), 29.1 (C5), 23.3 (C5-(CH₃)_{ax}), 22.1 (C5-(CH₃)_{eq}) and 15.9 (C2-CH₃); *m/z* 173 (5%, M⁺) and 129 (100, M – NMe₂); Anal. C₉H₁₉NO². Found: M⁺, 173.1407 requires *M*, 173.1416(Found: C, 62.30; H, 11.20; N, 7.90. C₉H₁₉NO₂ requires C, 62.40; H, 11.05; N, 8.10%). NOE results (Table 5) established the conformation (see above).

Table 5 Results of nOe experiments on **5**

| Signal irradiated | Percentage nOe detected | | | | | |
|-----------------------|-------------------------|------------------|--------------------|--------------------|-----------------------|-----------------------|
| | C(2)-Me | NMe ₂ | H(4) _{ax} | H(4) _{eq} | C(5-Me) _{eq} | C(5-Me) _{ax} |
| C(2)-Me | × | ? | – | – | – | – |
| NMe ₂ | 2.1 | × | 2.6 | – | – | – |
| H(4) _{ax} | – | 1 | × | 18 | 1.8 | – |
| H(4) _{eq} | – | – | 16 | × | 1.1 | 1.3 |
| C(5-Me) _{eq} | – | – | 3.2 | 2.0 | × | 0.6 |
| C(5-Me) _{ax} | – | – | – | 3.2 | 0.6 | × |

***N,N*,2-Trimethyl-1,3-dioxane-2-amine (6)**. A solution of 1,3-propanediol (0.988 g, 13 mmol) and a crystal of *p*-toluenesulphonic acid in dry toluene (20 mL) was heated to 100 °C under argon and *N,N*-dimethylacetamide dimethyl acetal (2.19 g, 16 mmol) added dropwise. The methanol-toluene azeotrope formed was distilled out and the solution stirred for 1 h at 100 °C. The solution was cooled, evaporated under reduced pressure and the residue distilled to give the *amide acetal* (0.84 g, 44 %) as a colorless oil, bp 44-46 °C / 12 mmHg; IR: $\nu_{\max}(\text{film})/\text{cm}^{-1}$ 2996, 2948 and 2878 (C–H), 2821 and 2783 (NMe₂), 1249, 1214, 1188, 1146, 1125, 1080 and 1050 (C–O); ¹H NMR (400 MHz; CDCl₃) δ 4.13 (2 H, ddd, *J* 12.7, 11.5 and 2.8, H_{4ax} and H_{6ax}), 3.57 (2 H, ddd, *J* 11.3, 5.2 and 1.3, H_{4eq} and H_{6eq}), 2.21 (6 H, s, -N(CH₃)₂), 2.01 (1 H, qt, *J* 12.9 and 5.2, H_{5ax}), 1.31 (1 H, dtt, *J* 13.2, 2.8 and 1.4, H_{5eq}) and 1.16 (3 H, s, C2-CH₃); ¹³C NMR (CDCl₃) δ 106.3 (C2), 58.4 (C4 and C6), 37.5 (-N(CH₃)₂), 24.6 (C5) and 16.4 (C2-CH₃); *m/z* 145 (100%, M⁺), 130 (25, M – Me) and 101 (100, M – NMe₂); HRMS: C₇H₁₅NO₂ Found M⁺, 145.1097, requires *M*, 145.1103; Anal. C₇H₁₅NO₂ Found: C, 58.00; H, 10.55; N, 9.50 requires C, 57.90; H 10.50; N, 9.65%).

***N,N*,1-Trimethylcyclohexylamine (7)**. 1-Methylcyclohexane carboxylic acid (4.85 g, 34 mmol) was added to freshly distilled thionyl chloride (40 mL) and heated under reflux for 2 h. The mixture was evaporated under reduced pressure and the residue distilled to give the acid chloride²² (4.26g, 78%) as a yellow oil, bp 80-81 °C/20 mmHg (lit. ²² 81 °C/20 mmHg); IR:

$\nu_{\max}(\text{film})/\text{cm}^{-1}$ 2937 and 2857 (C–H) and 1789 (C=O); $^1\text{H NMR}$ (200 MHz; CDCl_3): δ 2.17–1.23 (13 H, multiplet including 3 H singlet at 1.28); $^{13}\text{C NMR}(\text{CDCl}_3)$: δ 180.1 (–COCl), 53.7 (C1), 35.9 (C2 and C6), 25.8 (–CH₃), 25.3 (C4) and 22.8 (C3 and C5); MS: m/z 125 (80%, M – Cl). The acid chloride (4.4 g, 27.4 mmol) was dissolved in ether (50 mL), cooled to 0 °C and dry ammonia gas bubbled through for 2 h with stirring. The mixture was then washed with water (40 mL), the aqueous layer extracted with ether (2 × 30 mL) and the combined ether layers washed again with water (30 mL). The extract was dried (MgSO_4) and evaporated under reduced pressure to leave the amide²² (3.1 g, 79%) as a white solid which was used without further purification; mp 66–68 °C (lit.²² 68 °C); $R_f(\text{EtOAc})$ 0.40; IR: $\nu_{\max}(\text{KBr})/\text{cm}^{-1}$ 3435, 3401 and 3198 (N–H), 2919 and 2857 (C–H), 1649 (C=O) and 1612 (–CONH₂); $^1\text{H NMR}$ (200 MHz; CDCl_3) 6.06 (1 H, br s, –NH_AH_B), 5.70 (1 H, br s, –NH_AH_B), 1.91–1.25 (10 H, m) and 1.16 (3 H, s, –CH₃); $^{13}\text{C NMR } \delta\text{C}(\text{CDCl}_3)$ 180.9 (–CONH₂), 42.6 (C1), 35.6 (C2 and C6), 26.4 (–CH₃), 25.7 (C4) and 22.8 (C3 and C5); MS: m/z 141 (33%, M⁺), 126 (20, M – Me) and 97 (90, M – CONH₂); Found: M⁺, 141.1152. C₈H₁₅NO requires M , 141.1154.

Following the procedure of Hamlin and Freifelder,²³ the finely powdered 1-methylcyclohexanecarboxamide (2 g, 14.2 mmol) was added to a stirred solution of bromine (2.3 g, 14 mmol) in 20% aqueous potassium hydroxide (40 mL) at 0 °C. The solution was stirred for a further 30 min before extracting the isocyanate with ether (2 × 10 mL). The ether extract was carefully added dropwise to boiling concentrated hydrochloric acid (10 mL) and the mixture refluxed for 20 min. The mixture was then evaporated under reduced pressure and the residue recrystallized from EtOH–Et₂O to give

the amine hydrochloride (1.25 g, 60%), mp 195–200 °C dec.) (lit.²³ 285 °C dec.) The hydrochloride (1 g, 6.7 mmol) was dissolved in H₂O (10 mL), 20% aqueous KOH (10 mL) added and the mixture extracted with ether (3 × 20 mL). The extract was evaporated under reduced pressure and the residue distilled using a Kugelrohr apparatus to give the amine,²³ (0.45 g, 29%) as a colorless oil; IR: $\nu_{\max}(\text{Film})/\text{cm}^{-1}$ 3346 (N–H), 2925 and 2857 (C–H); $^1\text{H NMR}$ (200 MHz; CDCl_3): δ ca. 1.7 (2 H, br s, NH₂), 1.54–1.27 (10 H, m, (–CH₂)₅) and 1.05 (3 H, s, –CH₃); $^{13}\text{C NMR}$ (CDCl_3) 48.6 (C1), 40.7 (C2 and C6), 29.5 (–CH₃), 25.8 (C4) and 22.7 (C3 and C5); HRMS: m/z 113 (10%, M⁺), 98 (10, M – Me) and 70 (100, C₅H₁₀), Found: M⁺, 113.1204. C₇H₁₅N requires M , 113.1204. Finally, a mixture of the 1-methylcyclohexylamine (0.44 g, 3.8 mmol), formic acid (10 mL) and 40% aqueous formalin (0.75 mL) was refluxed for 4 h. The solution was evaporated under reduced pressure, 20 % aqueous NaOH (10 mL) added and then extracted with ether (3 × 10 mL). The extract was washed through a plug of silica with more ether and then evaporated under reduced pressure to give the amine²³ (0.16 g, 30%) as a colorless oil, IR: $\nu_{\max}(\text{Film})/\text{cm}^{-1}$ 2931 and 2815 (C–H) and 2776 (–N(CH₃)₂); $^1\text{H NMR}$ (250 MHz; CDCl_3): δ 2.19 (6 H, s, –N(CH₃)₂), 1.65–1.29 (10 H, m, (–CH₂)₅) and 0.84 (3 H, s, –CH₃); $^{13}\text{C NMR}$ (CDCl_3) δ 55.0 (C1), 37.5 (–N(CH₃)₂), 36.1 (C2 and C6), 26.0 (C4), 22.1 (C3 and C5) and 15.6 (–CH₃); HRMS: m/z 141 (80%, M⁺), 126 (50, M – Me) and 70 (90, C₅H₁₀), Found: M⁺ 141.1521. C₉H₁₉N requires M , 141.1517).

References

1. Kirby, A. J. *The Anomeric Effect and Related Stereoelectronic Effects at Oxygen*; Springer-Verlag: Berlin and Heidelberg, 1983.
2. (a) Carballeira, L.; Pérez-Juste, I. *J. Comput. Chem.* **2000**, *6*, 462. (b) Salzner, U.; Schleyer, P. von R. *Journal of Organic Chemistry* **1994**, *59*, 2138.
3. Jones, P. G.; Kirby, A. J.; Komarov, I. V.; Wothers, P. D. *J. Chem. Soc., Chem. Commun.* **1998**, 1695.
4. Kirby, A. J.; I V. Komarov; Wothers, P. D.; Feeder, N.; Jones, P. G. *Pure Appl. Chem.* **1999**, *71*, 385.
5. Costisella, V. B.; Gross, H. *J. Prakt. Chem.* **1977**, *319*, 8.
6. Lapuka, L. F.; Chalova, O. B.; Kantor, E. A.; Kiladze, T. K.; Rakhmankulov, D. L. *Khim. Geterotsikl. Soedin.* **1981**, 1182.
7. Pastushenko, E. V.; Lapuka, F. L.; Khatuntsev, I. I.; Zlot'skii, S. S.; Rakhmankulov, D. L. *Zh. Obshch. Khim.* **1986**, *56*, 59.
8. Allinger, N. L.; Graham, J. C. *J. Org. Chem.* **1971**, *36*, 1688.
9. Eliel, E. L.; Wilen, S. H. *Stereochemistry of Organic Compounds*; John Wiley & Sons: New York, 1994; p 696.
10. Booth, H.; Khedhair, K. A. *J. Chem. Soc., Chem. Commun.* **1985**, 467.
11. Eliel, E. L.; Knoebar, M. C. *J. Am. Chem. Soc.* **1968**, *90*, 3444.
12. Franck, R. W. *Tetrahedron* **1983**, *39*, 3251.
13. Graczyk, P. P.; Mikolajczyk, M. *Topics in Stereochemistry* **1994**, *21*, 209.
14. Booth, H.; Jozefowicz, M. L. *J. Chem. Soc., Perkin Trans. 2* **1976**, 895.
15. Höfner, D.; Lesko, S. A.; Binsch, G. *Org. Magn. Reson.* **1978**, *11*, 179.
16. Praly, J.-P.; Lemieux, R. U. *Can. J. Chem.* **1987**, *65*, 213.
17. Sanders, J. K. M., Personal Communication.
18. Booth, H.; Readshaw, S. A. *Tetrahedron* **1990**, *46*, 2097.
19. Booth, H.; Khedhair, K. A.; Readshaw, S. A. *Tetrahedron* **1987**, *43*, 4699.
20. de Hoog, A. J. *Org. Mag. Res.* **1974**, *6*, 233.
21. Arnold, Z.; Kornilov, M. J. *Collect. Czech. Chem. Commun.*, **1964**, *29*, 651.
22. Jacquier, R.; Christol, H. *Bull. Soc. Chim. Fr.* **1957**, 596.
23. Hamlin, K. E.; Freifelder, M. *J. Am. Chem. Soc.* **1953**, *75*, 369.